



Effect of magnesium carbonate on the uptake of aqueous zinc and lead ions by natural kaolinite and clinoptilolite

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Abstract

Adsorption behavior of Zn^{2+} and Pb^{2+} ions on kaolinite and clinoptilolite, originating from natural resources, was studied as a function of contact time and concentration. Zn^{2+} and Pb^{2+} ions are quickly adsorbed on both minerals and the uptake of the latter is more favored. The uptake of both ions was then examined on kaolinite– $MgCO_3$ and clinoptilolite– $MgCO_3$ mixtures over a metal ions range from 1 to 10000 mg/L. The sorption behavior of Zn^{2+} and Pb^{2+} on pure $MgCO_3$ was also studied. $MgCO_3$ is much more effective in the retention of Zn^{2+} and Pb^{2+} ions, in particular at higher concentrations. The large increase in the retarded amounts of both ions was associated with formation of the hydroxy-carbonate phases; namely hydrozincite for Zn^{2+} , and cerussite and hydrocerussite in the case of Pb^{2+} .

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1. Introduction

The bioavailability and mobility of naturally or anthropogenically released trace metal contaminants, such as Zn^{2+} and Pb^{2+} , in aquatic, aerial, and soil ecosystems are of great concern from the view point of environmental safety. Metal ions, which are natural constituents of the Earth's crust, can exert beneficial or harmful biological effects on the living organisms. Zn^{2+} ions are essential for humans, and deficiency

states with clinical abnormalities have been identified (Caussy et al., 2003). Zn^{2+} is commonly found in contaminated soils, sediments, and wastewater and groundwater streams. It accumulates in soils by atmospheric deposition originating from smelting operations and by agricultural applications of sewage sludge and agrochemicals and plays an essential role in cellular systems and enzymes (Nachtegaal and Sparks, 2004). Pb compounds are continuously discharged into the biological environments by various individual and industrial activities (gasoline, batteries, paints, insecticides, pipes, etc.). Due to its potential detriment to the living systems, Pb is classified at the top of the priority list of the most hazardous sub-

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stances along with As, Hg, C₂H₃Cl, and C₆H₆ (God-elitsas et al., 2003).

At the interface between aqueous media and soil, adsorption is an important process that can regulate the transport of chemical species through the hydrosphere. In industry, as far as the concentration of waste and the cost of treatment are considered, adsorption is an advantageous treatment method for removal of heavy metal ions from metal contaminated effluents. It was reported that sorption will be established into the aforementioned group of BATNEEC (best available technology not entailing excessive costs) treatment technologies as a standard process (Lehmann et al., 1999).

Kaolinite and clinoptilolite are often present as soil components. Due to the tight interlayer structure of kaolinite, most adsorption activities occurs along the edges and surfaces of the clay structure (Miranda-Trevino and Coles, 2003). Typically, kaolinite has a cation exchange capacity (CEC) of 3 to 15 meq/100 g, while values quoted for the specific surface area (SSA) of kaolinite are from 10 to 20 m²/g (Coles and Yong, 2002). Clinoptilolite is used in various applications, e.g. as ion exchanger, in gas separation, and removal of SO₂ and NH₃ (Tanaka et al., 2003). The CEC of clinoptilolite was reported to be 222–238 meq/100 g with Na⁺ or K⁺ as eluting cation, respectively (Langella et al., 2000).

Carbonate minerals are important constituents of soil and are frequently associated with other soil components like clay minerals, zeolites, and other mineral types, affecting the adsorption properties of these minerals. The presence of CaCO₃ in adsorbents was reported to correlate positively with the amount of removed metal ions from solutions (Wenming et al., 2001; Shahwan et al., 2002). MgCO₃ is among the most widely spread carbonate minerals in Earth's crust which plays an important role in regulating the aquatic environments as a result of precipitation, dissolution and sorption reactions. Based on a linear free energy correlation model, magnesite was predicted to have the largest trace metal partition coefficients among the carbonate minerals possessing a calcite structure (Wang and Xu, 2001).

The aim of this study is to examine the uptake behavior of Zn²⁺ and Pb²⁺ by kaolinite and clinoptilolite and to evaluate the effect of MgCO₃ on this uptake. So far, several studies were reported in litera-

ture regarding different aspects of the adsorption behavior of Zn²⁺ and Pb²⁺ by kaolinite and clinoptilolite (e.g. Mier et al., 2001; Mondale et al., 1995; Ouki and Kavannagh, 1997; Ponizovsky and Tsadilas, 2003; Coles and Yong, 2002; Langella et al., 2000; Ikhsan et al., 1999; Hooda and Alloway, 1998; Holm and Zhu, 1994). Our literature survey indicated a scarcity in the available data regarding the uptake of these ions by MgCO₃ (e.g. Lehmann et al., 1999).

2. Experimental

Kaolinite and clinoptilolite came from Sındırgı and Manisa regions located in the western part of Anatolia. The samples were dry-sieved and the particle size of kaolinite used in this study was 38–75 μm while that of clinoptilolite was 75–150 μm. The particle size of MgCO₃ (Fluka Chemika 63032) was 53–300 μm. Determination of the BET surface area was achieved by means of a Micrometrics ASAP 2010 instrument using nitrogen at 77 K. The samples were degassed for 3 h at 353 K.

CEC measurements were carried out by first exposing 1.0 g samples of kaolinite and clinoptilolite to 100 mL of fresh 0.5 M CaCl₂ solution on a daily basis under continuous mixing. This treatment was repeated three times, then the mineral samples were filtered and carefully washed. The adsorbed Ca²⁺ ions were successively leached using 100 mL aliquots of fresh 0.5 BaCl₂ solution, a process that was repeated four times on a daily basis. The released Ca²⁺ ions were then analyzed by AAS and ICP–AES.

The adsorption experiments were performed using 50-ml polyethylene tubes. To each tube, 0.50 g of kaolinite or clinoptilolite were added, followed by the addition of 50.0 ml of aqueous Pb(NO₃)₂ or Zn(NO₃)₂ solution. In the kinetic studies, the initial concentrations of the solutions were 1, 100, and 1000 mg/L and the contact periods were 10 min, 30 min, 2 h, 8 h, 24 h, and 48 h.

The concentrations of the solutions used to study the effect of MgCO₃ on the uptake by kaolinite and clinoptilolite were 1, 100, 500, 3000, and 10000 mg/L. Appropriate amounts of MgCO₃ were mixed with kaolinite or clinoptilolite to yield compositions of 5%, 10%, 25%, and 60% (w/w) of MgCO₃. The initial (at the start of mixing) and final (at the end of mixing) pH values obtained during the adsorption experiments are given in Tables 1 and 2. The speciation of both ions in aqueous media, in the absence of any complexing agent, and at two different concentrations (100 mg/L and 10000 mg/L) was analyzed using Visual MINTEQ software. This speciation model was also applied in earlier studies by other authors (e.g. Pierrard et al., 2002;

Table 1

Initial (at the start of mixing) and final (at the end of mixing) pH values of Zn^{2+} and Pb^{2+} solutions in contact with kaolinite, clinoptilolite, and $MgCO_3$

Cation (M^{2+})	Initial concentration $M(NO_3)_2$ (mg/L)	Kaolinite	Clinoptilolite	$MgCO_3$
Zn^{2+}	1.0	7.8–8.9	8.3–8.4	10.1–10.0
	100.0	6.4–7.1	6.7–7.4	8.1–9.6
	500.0	5.3–5.9	6.3–7.6	6.7–9.1
	3000.0	6.5–6.8	6.1–6.3	6.6–7.5
	10000.0	6.0–6.2	6.0–6.1	6.1–6.3
Pb^{2+}	1.0	7.7–8.0	8.6–8.0	9.9–9.8
	100.0	6.5–7.1	6.5–7.9	9.6–9.3
	500.0	5.3–5.9	6.3–7.6	6.3–9.0
	3000.0	4.8–5.4	4.1–5.0	4.9–8.9
	10000.0	4.1–4.7	3.9–4.4	4.1–7.7

Nour El-Dien et al., 1997). The data was generated based on parameters like initial concentration of ions, temperature, pH, and ionic strength, defined in an input file. The results indicated that the principal form of Pb at pH 1–7 is Pb^{2+} . As the pH increases beyond 8, hydroxyl species of Pb^{2+} are formed; $PbOH^+$ and Pb_2OH^{3+} , in addition to higher polycations. As the pH approaches 12, the major form becomes

$Pb(OH)_3^-$ together with $Pb(OH)_2$. In the case of Zn^{2+} , the speciation analysis revealed that Zn^{2+} ions exist up to pH of 8. At pH 8–11, the main form becomes $Zn(OH)_2$, and at pH > 11, $Zn(OH)_3^-$ dominates. Comparing these speciation with the pH data in Tables 1 and 2, the fraction of hydroxo species of both ions decreases when the initial concentration increases. In order to elucidate the effect of pH on the morphology of the evolving Zn- and Pb-carbonate species, separate sets of the experiments with initial concentrations of 10000 mg/L were performed at initial pH 10. The pH was adjusted by a 0.1 M NaOH solution.

The experiments were conducted using the batch method at 25 °C under atmospheric pressure. The dispersions were shaken for 48 h (Nuve ST 402 water bath shaker equipped with a microprocessor thermostat). At the end of each mixing period, the samples were filtered and dried at 90 °C. The filtrates were analyzed by flame AAS using a Thermo Elemental SOLAAR M6 Series atomic absorption spectrometer (Pb and Zn hollow cathode lamps) with air–acetylene flame.

XRPD was performed using a Philips X'Pert Pro diffractometer (Cu K_{α} radiation, step size of 0.020 2 θ). SEM/EDS characterization was carried out with a Philips XL-30S FEG type instrument. The solid samples were sprinkled onto adhesive carbon tapes supported on metallic disks. EDS

Table 2

Initial (at the start of mixing) and final (at the end of mixing) pH values of Zn^{2+} and Pb^{2+} solutions in contact with kaolinite– $MgCO_3$, and clinoptilolite– $MgCO_3$ mixtures

Cation	Initial concentrate (mg/L)	Mixture	$MgCO_3$ (wt.%)			
			5	10	25	60
Zn^{2+}	1.0	Kaolinite– $MgCO_3$	9.5–9.7	9.6–9.7	10.1–9.3	10.1–9.8
	100.0		7.5–9.1	7.4–9.4	7.2–9.4	7.7–9.4
	500.0		6.8–9.3	6.8–9.5	6.9–9.4	7.2–9.6
	3000.0		6.5–6.8	6.5–5.7	6.6–6.8	6.6–8.7
	10000.0		6.0–6.1	6.1–5.9	6.1–6.1	6.1–7.1
Pb^{2+}	1.0	Kaolinite– $MgCO_3$	9.6–9.2	9.3–9.2	9.8–9.4	9.7–9.7
	100.0		6.5–8.5	6.6–8.2	8.2–8.8	9.1–9.4
	500.0		5.3–9.3	5.7–9.5	5.8–9.4	6.1–9.6
	3000.0		5.0–5.6	5.0–6.6	4.9–8.7	4.9–8.9
	10000.0		4.2–4.6	4.0–4.6	4.0–4.9	4.0–5.1
Zn^{2+}	1.0	Clinoptilolite– $MgCO_3$	9.5–9.5	9.4–9.5	9.6–9.3	9.9–9.4
	100.0		7.2–9.1	7.8–9.3	8.7–9.3	8.1–9.2
	500.0		6.3–9.1	6.4–9.1	6.2–9.6	6.6–9.2
	3000.0		6.3–6.4	6.4–6.6	6.3–8.3	6.3–8.3
	10000.0		6.1–5.9	6.1–5.9	6.2–6.1	6.1–7.2
Pb^{2+}	1.0	Clinoptilolite– $MgCO_3$	9.4–8.8	9.3–9.1	9.1–8.9	9.2–9.2
	100.0		7.8–9.1	7.6–9.1	9.1–8.7	9.6–9.1
	500.0		6.3–9.1	6.4–9.1	6.2–9.6	6.0–9.2
	3000.0		4.4–5.1	4.5–8.4	4.1–8.8	5.6–9.2
	10000.0		3.9–4.6	3.8–4.8	3.8–5.4	3.7–8.5

In all cases, M/V ratio is 0.5 g solid/50 mL solution.

elemental analysis was performed at different points on the surface in order to minimize possible anomalies arising from the heterogeneous nature of the analyzed surface.

3. Results and discussion

The mineralogical purity of the natural samples of kaolinite and clinoptilolite was tested using XRPD. Kaolinite samples were composed of kaolinite in addition to quartz as the main impurity whereas clinoptilolite appeared to be almost pure. EDS indicated that the average elemental content of kaolinite was 66.8% O, 17.6% Si, 14.1% Al in addition to minor quantities of Na, K, Mg, and Ca probably originating from a non-kaolinitic impurity that is below the detection limit of XRPD. The average elemental composition of clinoptilolite was 61.9% O, 23.7% Si, 5.3% Al, including small amounts of Na, K, Mg, and Ca. For MgCO_3 , the average elemental content was 15.6% Mg, 21.2% C, and 61.3% O in addition to a minor amount of Ca. The EDS values given above represent an average of five data points obtained from random locations on the mineral surface, and correspond to the upper surface ($\sim 2 \mu\text{m}$ depth). SEM images indicated that the kaolinite particles have well-defined hexagonal plates with edge sizes from 300 to 500 nm. Clinoptilolite seems to be composed of crystals with varying shape and size of several micrometers. MgCO_3 was composed mostly of aggregates with variable shapes and sizes and lacked a distinct morphology.

The CEC values for kaolinite and clinoptilolite samples were 21.4 and 192.6 meq $\text{Ca}^{2+}/100 \text{ g}$ of the mineral. The CEC of the kaolinite applied in our studies appears to be larger than the typical values reported for pure kaolinite being usually within 3–15 meq/100 g range (Coles and Yong, 2002). This is possibly attributed to the presence of smectite-like impurities, which could enhance the CEC significantly even if present in small amounts. The presence of such an impurity could be responsible for the presence of the exchangeable elements; Na, K, Mg, and Ca – mentioned above – in natural kaolinite. On the other hand, the CEC of clinoptilolite applied in this work seems to be somewhat smaller than typical values reported to vary within 222–238 meq/100 g (Langella et al., 2000). The specific surface area

(BET) of kaolinite is $6.9 \text{ m}^2/\text{g}$, of clinoptilolite is $31.7 \text{ m}^2/\text{g}$. Surface area determination by BET– N_2 was reported to yield lower values as compared to EGME method (Coles and Yong, 2002).

The effect of time of mixing on the adsorption of Pb^{2+} and Zn^{2+} kaolinite was studied at room temperature at the initial concentrations of 1, 100, and 1000 mg/L of the nitrate solutions of both ions. Equilibrium was reached faster at the initial concentrations of 1 and 100 mg/L than at the 1000 mg/L. Adsorption was faster for kaolinite than for clinoptilolite. The adsorption sites on kaolinite are located mainly on the surface and edge sites of the clay mineral and are more readily accessible than those of clinoptilolite, an important fraction of which is provided by the channels within the structure. Over the entire range of initial concentrations, equilibrium is generally achieved within the first 24 h of mixing. The further experiments were carried out at this mixing period.

The effect of concentration on the uptake of aqueous Zn^{2+} and Pb^{2+} ions by kaolinite and clinoptilolite was investigated at the initial concentrations of $\text{Zn}(\text{NO}_3)_2$ and $\text{Pb}(\text{NO}_3)_2$ of 1, 100, 500, 3000, and 10000 mg/L. The results indicated that almost a complete removal of both ions was achieved by both minerals for the initial concentrations of 1 and 100 mg/L. The adsorption data of both ions on kaolinite and clinoptilolite were examined using the Freundlich isotherm; the linearized form of which is given by:

$$\log[C]_s = \log k + n \log[C]_l$$

where $[C]_s$ is the amount adsorbed, $[C]_l$ is the equilibrium concentration of the adsorptive (mg/L), and n and k are the Freundlich constants. The adsorption isotherms of Zn^{2+} on kaolinite and clinoptilolite were reported before (Shahwan et al., 2005) and are provided here for comparison with those of Pb^{2+} ions. Freundlich plots are given in Fig. 1 and the values of n and k are provided in Table 3. The relatively low n values imply that, within the studied concentration range, increasing the initial concentration is resulting in a significant decrease in the fixation ability of the minerals. The k values indicate a higher adsorption capacity of clinoptilolite compared to kaolinite towards both

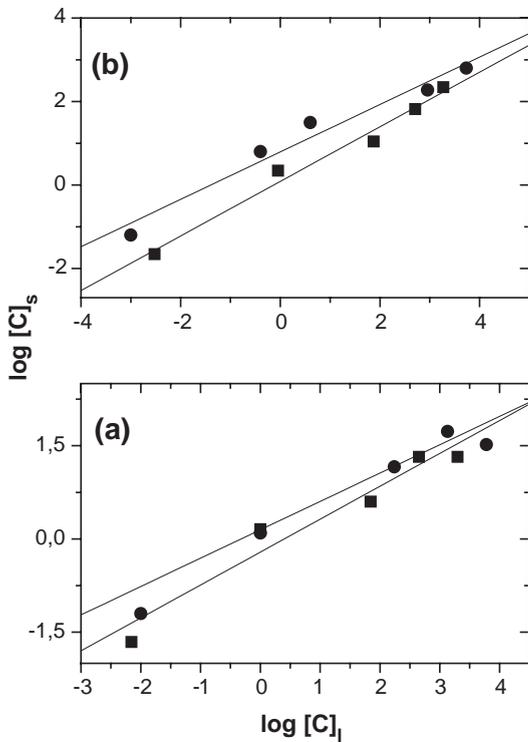


Fig. 1. Freundlich plots corresponding to the adsorption of (a) Zn²⁺ ions and (b) Pb²⁺ ions on kaolinite and clinoptilolite: (■) kaolinite and (●) clinoptilolite.

ions, and reveals also higher capacity of both minerals towards Pb²⁺.

The XRPD and SEM characterization showed no changes in the structure and particle morphology of kaolinite and clinoptilolite upon the sorption of Zn²⁺ and Pb²⁺ by the two minerals. No surface precipitation was also detected, which was expected as a result of pH 4–8 (Table 1).

Cation exchange is the primary retention mechanism for Pb²⁺ by kaolinite and the speciation of this ion

Table 3
Freundlich parameters for the adsorption of Zn²⁺ and Pb²⁺ ions on kaolinite and clinoptilolite

	<i>n</i>	<i>k</i>	<i>R</i>
Zn–kaolinite	0.53	0.62	0.9583
Zn–clinoptilolite	0.65	1.21	0.9908
Pb–kaolinite	0.46	1.41	0.9295
Pb–clinoptilolite	0.57	6.21	0.9849

in solution could affect the selectivity of the kaolinite towards the same ion (Coles and Yong, 2002). In another study, the uptake of Pb²⁺ ions by kaolinite was referred to a combination of ion exchange and surface complexation, the hypothesis supported by the agreement between an ion exchange/surface complexation model and the sorption data (Holm and Zhu, 1994). In a study of Pb²⁺ and Zn²⁺ retention by kaolinite at metal concentrations of 100 μM, it was shown that the adsorption data were fitted closely by a constant–capacitance surface complexation model, one involving electrostatic attraction between transition metal ions and the permanent, negatively charged sites on the silanol faces of kaolinite, while the other is an inner-sphere complex at the variable-charge surface hydroxyl groups situated at the crystal edges and on the aluminol faces (Ikhsan et al., 1999). Our previous studies on the adsorption of other divalent ions (Co²⁺ and Ba²⁺) by kaolinite using ToF–SIMS depth profiling have shown that the fixation of these ions is associated with a decrease in the initial amount of exchangeable cations of the kaolinite, but that the

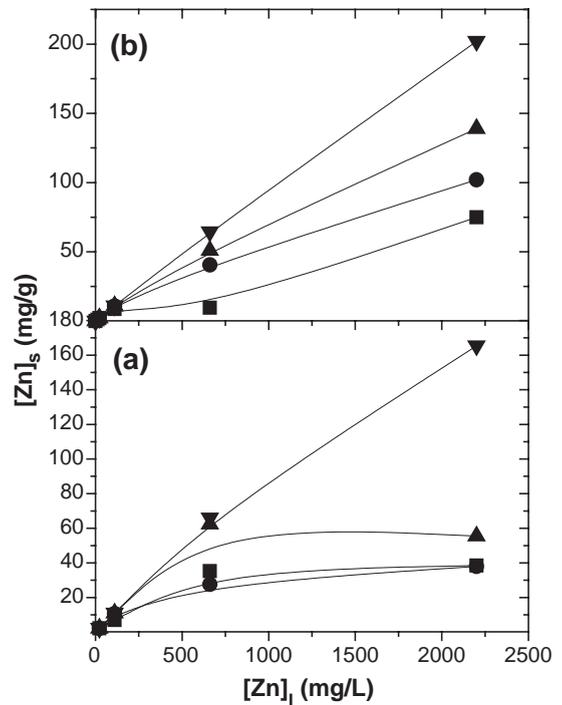


Fig. 2. Adsorption isotherms of Zn²⁺ ions on (a) kaolinite–MgCO₃ and (b) clinoptilolite–MgCO₃ mixtures: (■) 5% MgCO₃, (●) 10% MgCO₃, (▲) 25% MgCO₃ and (▼) 60% MgCO₃.

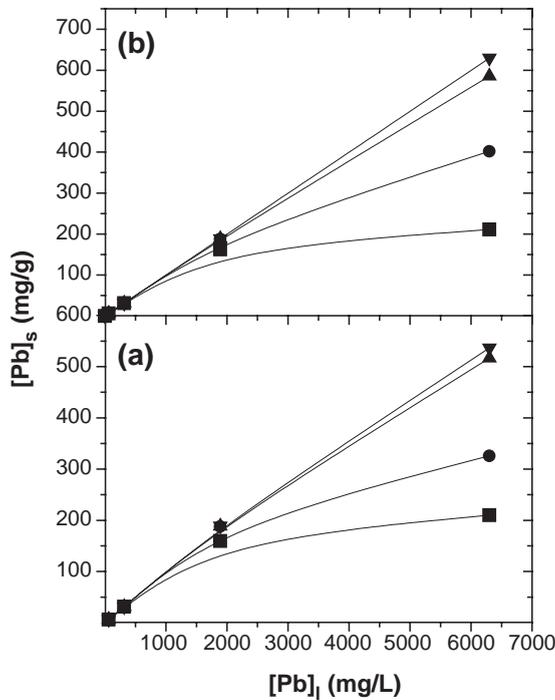


Fig. 3. Adsorption isotherms of Pb^{2+} ions on (a) kaolinite– MgCO_3 and (b) clinoptilolite– MgCO_3 mixtures: (■) 5% MgCO_3 , (●) 10% MgCO_3 , (▲) 25% MgCO_3 and (▼) 60% MgCO_3 .

decrease did not quantitatively agree with the amount of adsorbed Co^{2+} or Ba^{2+} kaolinite (Shahwan et al., 2004). In summary, when discussing the adsorption mechanism by kaolinite, in addition to the exchange of exchangeable metal cations, protons of the OH_2^+ and OH groups at the edges should also be taken into consideration. Usually the contribution of each mechanism depends on the pH conditions, crystallinity, and particle size of the sample (Ma and Eggleton, 1999).

Due to its porous structure that usually contains readily exchangeable ions like Na^+ , K^+ , and Ca^{2+} , clinoptilolite is expected to adsorb metal ions mainly through ion exchange as reported for the adsorption of Zn^{2+} and Pb^{2+} in a number of studies (e.g. Ponizovsky and Tsadilas, 2003; Ouki and Kavannagh, 1997; Mondale et al., 1995). Higher selectivity was reported for Pb^{2+} by clinoptilolite compared to Zn^{2+} (Langella et al., 2000; Mondale et al., 1995), in line with our findings. The ion-exchange mechanism is expected to prevail unless the pH is high enough to allow possible precipitation of those ions as hydroxides.

High pH might also lead to the formation of anionic hydroxo-complexes of the metal ions, thus lowering the adsorption (Mier et al., 2001). It should be noted that different cationic compositions can lead to contrasting behavior of various clinoptilolites towards metal ions, ascribed to different characteristics of the deposits of this mineral (Langella et al., 2000). Generally, the wide variation in the cation exchange capacity of zeolites was reported to be caused by the differing nature of various zeolite cage structures, natural structural defects, adsorbed ions, and their associated gangue minerals (Mondale et al., 1995). The same study documented also a high selectivity of clinoptilolite towards Pb^{2+} .

The amount of adsorbed Zn^{2+} and Pb^{2+} ions increases with the increasing MgCO_3 in the mixtures with kaolinite and clinoptilolite (Figs. 2 and 3). In all cases, the amounts of both cations adsorbed seem to vary almost linearly with increasing equilibrium concentration – within the studied

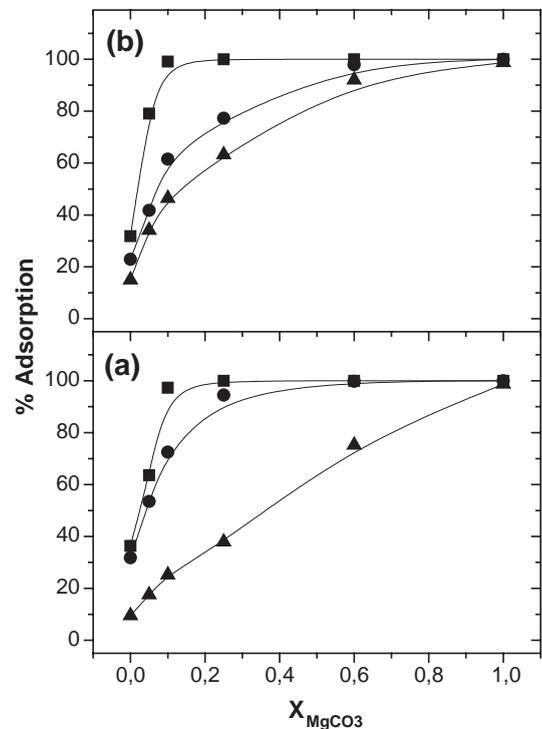


Fig. 4. Adsorption of Zn^{2+} ions (in percent added amounts) with the mass fraction of MgCO_3 : (a) kaolinite– MgCO_3 and (b) clinoptilolite– MgCO_3 . Initial concentrations are (■) 500 mg/L, (●) 3000 mg/L, and (▲) 10000 mg/L of $\text{Zn}(\text{NO}_3)_2$.

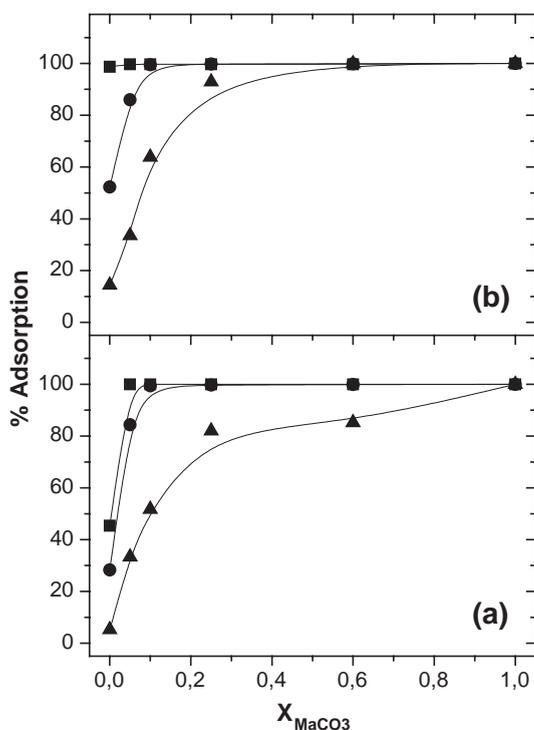


Fig. 5. Adsorption of Pb²⁺ (in percent added amounts) ions with the mass fraction of MgCO₃: (a) kaolinite–MgCO₃ and (b) clinoptilolite–MgCO₃. Initial concentrations are (■) 500 mg/L, (●) 3000 mg/L, and (▲) 10000 mg/L of Pb(NO₃)₂ solutions.

range of concentrations – as the MgCO₃ fraction was increased to 0.60. The increase of the adsorption expressed in % of metal ions added is shown in Figs. 4 and 5. The figures display a large increase in the efficiency of removal of both ions as the fraction of MgCO₃ in the mixtures was increased.

The XRPD analysis showed that the large increase in the amounts of Zn²⁺ and Pb²⁺ ions adsorbed by kaolinite–MgCO₃ or clinoptilolite–MgCO₃ at higher concentrations was associated with the precipitation of hydroxo-carbonate phases of these ions, in particular at high MgCO₃ contents. The percentage uptake on MgCO₃ was drawn as a function of initial concentration of Zn(NO₃)₂ and Pb(NO₃)₂ solutions (Fig. 6), which gives also the corresponding behavior of kaolinite and clinoptilolite. The efficiency of MgCO₃ in the removal of both cations is more distinct at higher initial concentration, where kaolinite and clinoptilolite becomes satu-

rated. XRPD analysis of Zn²⁺- and Pb²⁺-loaded MgCO₃ showed that the principal mechanism of retardation of Zn²⁺ and Pb²⁺ ions at the highest initial concentrations (3000 and 10000 mg/L) is mainly through precipitation. At the lower initial concentration (1, 100 and 500 mg/L) no precipitates were detected for both ions, possibly indicating that the sorption process is dominated by ion-exchange. In addition to the enhancement of removal of ions from solution, precipitation is reported to greatly immobilize the adsorbed ions (Wang and Xu, 2001). Zn²⁺-loaded MgCO₃ samples revealed the formation of hydrozincite (Fig. 7a). In the presence of carbonate and hydroxide ions, zinc ions can form precipitates of the general formula Zn(CO₃)_x(OH)_{2–2x} with 0 < x < 1, where x is controlled by the synthesis procedure (Zhang et al., 2004). The formation of hydrozincite in our experiments is interesting because the pH of the medium at the initial concentrations of 3000 and 10000 mg/L (at which the precipitates occurred) was slightly acidic (Table 1). This indicates that the solubility limit is exceeded

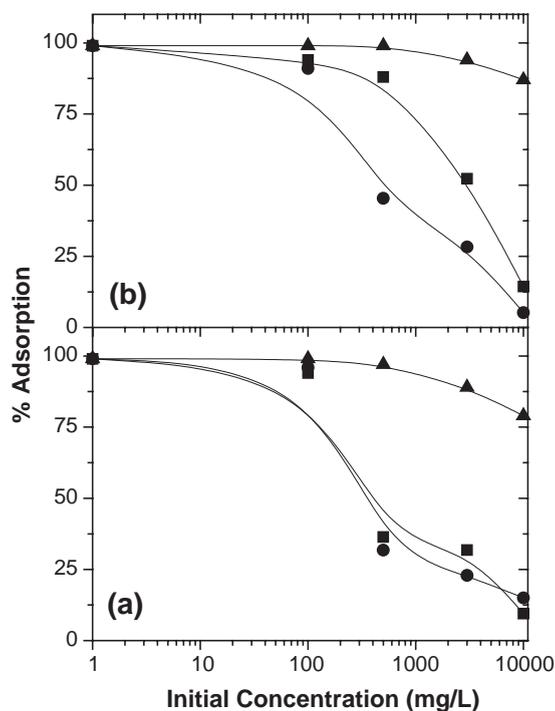


Fig. 6. Adsorption of (a) Zn ions and (b) Pb ions (in percent added amount) on kaolinite (●), clinoptilolite (■), and magnesite (▲) as a function of initial metal ion concentration.

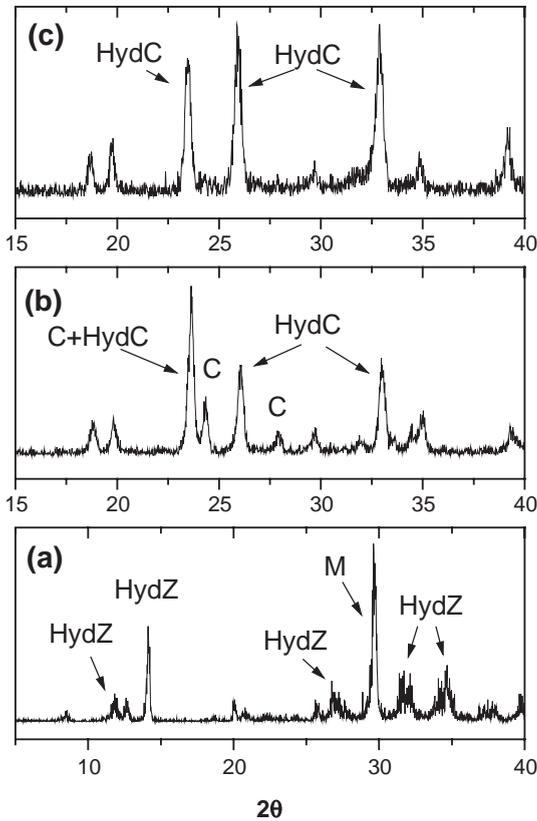
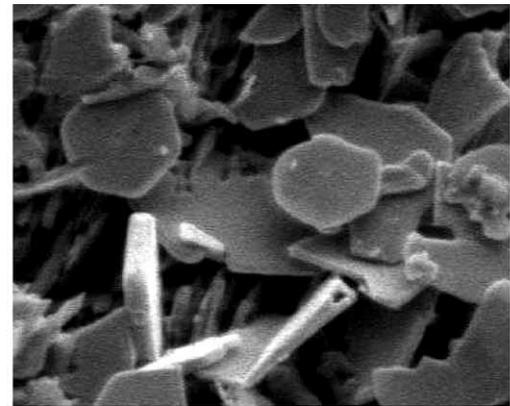


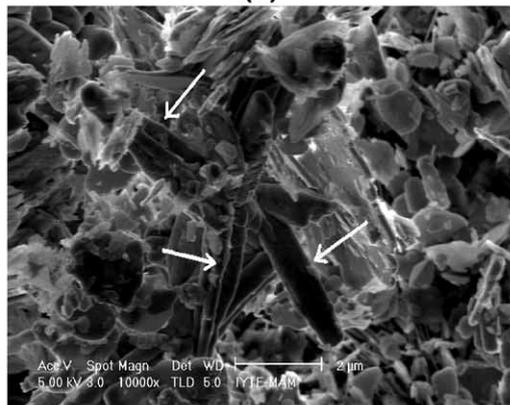
Fig. 7. XRPD patterns showing (a) hydrozincite formation on MgCO_3 in the pH range of 6.1–6.3, (b) cerussite and hydrocerussite formation on MgCO_3 in the pH range of 4.1–7.7 and (c) hydrocerussite formation on MgCO_3 when the initial pH is 10. In all cases, the initial concentration of $\text{Zn}(\text{NO}_3)_2$ and $\text{Pb}(\text{NO}_3)_2$ is 10000 mg/L. HydZ: hydrozincite; HydC: hydrocerussite; C: cerussite.

near the carbonate surface due to the accumulation of Zn^{2+} ions in the diffuse part of the electrical double layer, thus allowing for surface precipitation. XRPD and SEM analysis showed that increasing the initial pH to 10 did not change the type of precipitated species although more amount of precipitate was observed. A typical SEM image of hydrozincite is shown in Fig. 8a.

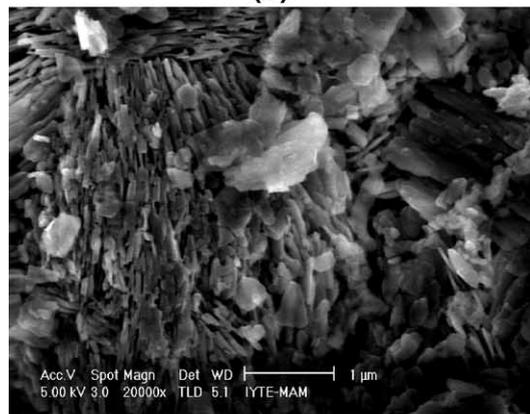
For the Pb-loaded MgCO_3 samples, Fig. 7b indicates the formation of a mixture of cerussite, PbCO_3 , and hydrocerussite, $\text{Pb}_3(\text{CO}_3)_2(\text{OH})_2$, in the samples that were prepared without any pH control. Again, the partial presence of hydrocerussite is interesting as the pH of the medium at the corresponding initial concentrations of Pb^{2+} solutions was acidic (Table



(c)



(b)



(a)

Fig. 8. Typical SEM images of (a) hydrozincite, (b) cerussite crystals embedded between crystals of hydrocerussite and MgCO_3 and (c) hydrocerussite crystals.

1). Increasing the initial pH to 10.0 lead to the formation of hydrocerussite only (Figs. 7c and 8c), indicating that this form is more stable at high pH values. SEM studies supported the XRPD observations. Cerussite crystals possess a columnar morphology (follow the arrows; approximate length of crystals is 5 μm ; approximate width is 0.5 μm) embedded in hydrocerussite crystals. Hydrocerussite crystals have plate-like morphology with shapes close to hexagonal but with rounded edges, and edge lengths ranging from 0.5 to 2 μm .

Based on a linear free energy correlation model, magnesite among isostructural carbonates was reported to possess the largest partition coefficients, K_d , for divalent metal ions (Wang and Xu, 2001). The model, which was based on a theoretical thermodynamic approach, considered that metal ion partitioning between minerals and aqueous solution depends on the bonding energies between the fixed metal ion and the host mineral, the size difference between the fixed metal and the host cation, and the difference in chemical potential between the fixed metal cation and the host one in solution. Based on these considerations, $\log K_d$ values of Zn^{2+} and Pb^{2+} on magnesite were predicted to be 1.97 and 4.26, respectively. Together with stressing the importance of MgCO_3 as a scavenger for metal ions, these theoretical predictions are qualitatively in agreement with our results which showed that larger amounts of Pb^{2+} ions are fixed by MgCO_3 compared to Zn^{2+} (Fig. 6).

4. Conclusions

The adsorption of Zn^{2+} and Pb^{2+} ions on kaolinite and clinoptilolite is fast and obeys the Freundlich isotherm model. The affinity of both minerals towards Pb^{2+} ions is generally larger than towards Zn^{2+} ions. Within the studied concentration range and pH conditions, the presence of MgCO_3 increased the adsorption of Zn^{2+} and Pb^{2+} ions. The large increase in the amounts of Zn^{2+} and Pb^{2+} ions adsorbed with the increasing fraction of MgCO_3 reveals that MgCO_3 is a more favorable sink for these ions and demonstrates the higher uptake capacity of this mineral compared to kaolinite and clinoptilolite. At higher fractions of MgCO_3 and higher equilibrium concentrations, pre-

cipitation of hydroxo-carbonates enhances the removal of both ions from the aqueous solutions.

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